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Oxidation And Reduction Practice Problems Answers

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Oxidation And Reduction Practice Problems

Practice: Redox reactions questions. This is the currently selected item.

Oxidizing and reducing agents.

Disproportionation.

Balancing redox

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reactions in acid.

Balancing redox

reactions in base.

Answers

Redox reactions questions (practice)

| Khan Academy

Practice Problems

Oxidation & Reduction.

1. Some anaerobic bacteria utilize oxidizing agents other than O_2 as an energy source; for example, SO_4^{2-} , NO_3^- , and Fe^{3+} . One half-reaction is $FeO(OH)(s) + HCO_3^-$

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$\text{Fe}^{2+}(\text{aq}) + 2 \text{H}^{+}(\text{aq}) + \text{e}^{-}$
 $\rightarrow \text{FeCO}_3(\text{s}) + 2 \text{H}_2\text{O}(\text{l})$, for which $E^{\circ} =$
 $+1.67 \text{ V}$. What mass of
iron gives the same
standard reaction
Gibbs energy as 1.00 g
of oxygen?

Oxidation/Reduction Practice Problems

Practice Problems:

Redox Reactions.

Determine the
oxidation number of
the elements in each of
the following

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compounds: a. H_2CO_3

b. N_2 c. Zn(OH)_2

d. NO_2 e. LiH f. Fe_3O_4

Hint; Identify the species being oxidized and reduced in each of the following reactions:

a. $\text{Cr} + \text{Sn}^{4+} \rightarrow \text{Cr}^{3+} + \text{Sn}^{2+}$

b. $3\text{Hg}^{2+} + 2\text{Fe(s)} \rightarrow 3\text{Hg} + 2\text{Fe}^{3+}$

c. $2\text{As(s)} + 3\text{Cl}_2(\text{g}) \rightarrow 2\text{AsCl}_3$

Hint

Hint

Practice Problems:

Redox Reactions

Oxidation-Reduction

Balancing Additional

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Practice Problems

Acidic Solution 1. $\text{Ag} + \text{NO}_3^- \rightarrow \text{Ag}^+ + \text{NO}$

Answer: $4\text{H}^+ + 3\text{Ag} + \text{NO}_3^- \rightarrow 3\text{Ag}^+ + \text{NO} + 2\text{H}_2\text{O}$

2. $\text{Zn} + \text{NO}_3^- \rightarrow \text{Zn}^{2+} + \text{NH}_4^+$

Answer: $10\text{H}^+ + 4\text{Zn} + \text{NO}_3^- \rightarrow 4\text{Zn}^{2+} + \text{NH}_4^+ + 3\text{H}_2\text{O}$

3. $\text{Cr}_2\text{O}_7^{2-} + \text{C}_2\text{H}_4\text{O} \rightarrow \text{C}_2\text{H}_4\text{O}_2 + \text{Cr}^{3+}$

Answer: $8\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 3\text{C}_2\text{H}_4\text{O} \rightarrow 3\text{C}_2\text{H}_4\text{O}_2 + 2\text{Cr}^{3+} + 4\text{H}_2\text{O}$

4. $\text{H}_3\text{PO}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{H}_3\text{PO}_4 + \text{Cr}^{3+}$

Answer: $8\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 3\text{H}_3\text{PO}_2 \rightarrow 3\text{H}_3\text{PO}_4 + 2\text{Cr}^{3+} + 4\text{H}_2\text{O}$

5. $\text{Cr}_2\text{O}_7^{2-} + \text{C}_2\text{H}_4\text{O} \rightarrow \text{C}_2\text{H}_4\text{O}_2 + \text{Cr}^{3+}$

Answer: $8\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 3\text{C}_2\text{H}_4\text{O} \rightarrow 3\text{C}_2\text{H}_4\text{O}_2 + 2\text{Cr}^{3+} + 4\text{H}_2\text{O}$

6. $\text{H}_3\text{PO}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{H}_3\text{PO}_4 + \text{Cr}^{3+}$

Answer: $8\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 3\text{H}_3\text{PO}_2 \rightarrow 3\text{H}_3\text{PO}_4 + 2\text{Cr}^{3+} + 4\text{H}_2\text{O}$

7. $\text{H}_3\text{PO}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{H}_3\text{PO}_4 + \text{Cr}^{3+}$

Answer: $8\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 3\text{H}_3\text{PO}_2 \rightarrow 3\text{H}_3\text{PO}_4 + 2\text{Cr}^{3+} + 4\text{H}_2\text{O}$

8. $\text{H}_3\text{PO}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{H}_3\text{PO}_4 + \text{Cr}^{3+}$

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Oxidation-Reduction

Extra Practice Problems

ScienceGeek.net

In an oxidation-reduction or redox reaction, it is often confusing to identify which molecule is oxidized in the reaction and which molecule is reduced. This example problem shows how to correctly identify which atoms undergo oxidation or reduction and their corresponding redox

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agents.

Practice Problems

Oxidation and

Reduction Reaction

Example Problem

Practice Problems:

Redox Reactions

(Answer Key)

Determine the

oxidation number of

the elements in each of

the following

compounds: a. H_2CO_3

H: +1, O: -2, C: +4 b.

N_2 N: 0 c. $\text{Zn}(\text{OH})_2$

Zn: 2+, H: +1, O: -2

d. NO_2 N: +3, O: -2 e.

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LiH Li: +1, H: -1 f. Fe₃

O₄ Fe: +8/3, O: -2;

Identify the species

being oxidized and

reduced in each of the

...

Practice Problems: Redox Reactions

This worksheet and
quiz let you practice
the following skills:

Reading

comprehension -

ensure that you draw

the most important

information from

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**Quiz & Worksheet -
Oxidation &
Reduction Reactions
in the ...**

Oxidation-Reduction reactions (also called “redox” reactions) are reactions that involve a shift of electrons between reactants. Oxidation is complete or partial loss of electrons or gain of

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oxygen. The loss of electrons results in an increase in charge or oxidation state.

Reduction is complete or partial gain of electrons or loss of oxygen.

Oxidation-Reduction Reactions Quiz - Softschools.com

Problem #8: $\text{Fe} + \text{HCl} \rightarrow \text{HFeCl}_4 + \text{H}_2$.

Solution: 1) This problem poses interesting problems,

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especially with the Cl.

The key to solving this problem is to eliminate everything not directly involved in the redox.

That means the H in HFeCl_4 as well as the Cl in it and HCl. When we do that, this is the unbalanced, ionic form we wind up with:

Balancing redox reactions in acidic solution: Problems #1-10

Examples of oxidation

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reduction (redox)
reactions, oxidizing

and reducing agents,
and common types of

redox reactions. If
you're seeing this
message, it means
we're having trouble
loading external
resources on our
website.

Oxidation-reduction

(redox) reactions

(article) | Khan

Academy

Reduction occurs when

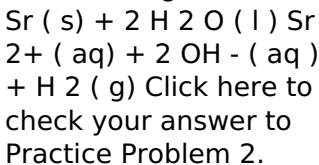
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the oxidation number of an atom becomes smaller. Practice

Problem 2: Determine which atom is oxidized and which is reduced in the following reaction.



Oxidation and Reduction - Purdue University

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Oxidation & Reduction.

1. Some anaerobic

bacteria utilize

oxidizing agents other

than O_2 as an energy

source; for example,

SO_4^{2-} , NO_3^- , and Fe^{3+} .

One half-reaction

is $FeO(OH)(s) + HCO_3^-$

$(aq) + 2 H^+(aq) + e^-$

$\rightarrow FeCO_3(s) + 2 H_2O(l)$,

for which $E^\circ =$

$+1.67 V$. What mass of

iron gives the same

standard reaction

Gibbs energy as 1.00 g

of oxygen?

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Oxidation/Reduction Practice Problems Answers

You also know that oxidation and reduction reactions occur in pairs: if one species is oxidized, another must be reduced at the same time - thus the term 'redox reaction'. Most of the redox reactions you have seen previously in general chemistry probably

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involved the flow of electrons from one metal to another, such as the reaction between ...

10.10: Oxidation and Reduction in Organic Chemistry ...

B. reduction, only C. both oxidation and reduction D. neither oxidation nor reduction

23. In the reaction $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{AgCl}(\text{s})$, the reactants A. gain

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Practice Problems

Answers

electrons, only B. lose electrons, only C. both gain and lose electrons D. neither gain nor lose electrons 24. In the reaction $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$, the correct half-reaction for the ...

Redox practice worksheet

Oxidation is the loss of electrons or an increase in oxidation state by a molecule, atom, or ion. Reduction

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Oxidation And

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Practice Problems

is the gain of electrons or a decrease in oxidation state by a molecule, atom, or ion.

As an example, during the combustion of wood, oxygen from the air is reduced, gaining electrons from carbon which is oxidized.

Oxidation and Reduction | Exams Daily

Answer: The hydrogen atoms have an oxidation state of +1

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Practice Problems

Answers

and the oxygen atom has an oxidation state of -2. Problem: Assign oxidation states to each atom in CaF_2 . Calcium is a Group 2 metal. Group IIA metals have an oxidation of +2. Fluorine is a halogen or Group VIIA element and has a higher electronegativity than calcium. According to rule 8 ...

Assigning Oxidation

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Oxidation And

Reduction

States Example

Problem

PreAP Chemistry

-REDOX Practice

Problems Directions:

Identify with substance
is undergoing oxidation

and reduction. Also,

Identify the Oxidizing

Agent and the

Reducing Agent in

each equation.

$\text{HNO}_3(\text{aq}) +$

$\text{H}_3\text{AsO}_3(\text{aq}) \rightarrow \text{NO}(\text{g}) +$

$\text{H}_3\text{AsO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l})$

IB Chemistry SL -

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Oxidation And

Reduction

Additional REDOX Practice Problems

Balancing REDOX

Reactions: Learn and

Practice Reduction-

Oxidation reactions (or
REDOX reactions)

occur when the

chemical species

involved in the

reactions gain and lose

electrons. Oxidation

and reduction occur

simultaneously in order

to conserve charge. We

can “see” these

changes if we assign

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Oxidation And

Reduction

oxidation numbers to

the reactants and

products.

Practice Problems
Answers

Balancing REDOX Reactions: Learn and Practice

Just like for the

alkenes, the OH group

is placed on the less

substituted

carbon. However, this

time an enol is formed

which quickly

rearranges into the

corresponding carbonyl

via keto-enol

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Problems

tautomerization.. The reaction is regioselective when a terminal alkyne is used. In this case, an aldehyde is formed. For hydroboration-oxidation of alkynes, the use of alkyl boranes is preferred, especially ...

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