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ISOTOPES AND

AVERAGE Name

ATOMIC MASS

Elements come in a variety of isotopes, meaning they are made up of atoms with the same atomic number but different atomic masses These atoms differ in the number

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BONDS Name Classify  
the following

compounds as ionic  
(metal + nonmetal),  
covalent (nonmetal +

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Name NAMING

MOLECULAR

COMPOUNDS Name the  
following covalent

compounds. 12.. 16.

17. 19. 20. So Cac IGde  
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5. 35 g of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$   
Determine the number  
of grams in each of the  
quantities below. 1, 2.5  
moles of  $\text{NaCl}$  2. 0.50  
moles of 3, 1.70 moles  
of  $\text{KMnO}_4$  4, 0.25 moles  
of  $\text{KCl}$  e, ST 5. 3.2  
moles of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$   
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A liquid will boil when  
its vapor pressure  
equals atmospheric  
pressure. Answer the  
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LE CHATELIER'S

PRINCIPLE Name \_ Le

Chatelier's Principle

states that when a

system at equilibrium

is subjected to a stress,

the system will shift its

equilibrium point in

order to relieve the

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TABLE CHAPTER 6 & 7

WORKSHEET Read the

clue and give the

element symbol that

applies. CLUE 1. 2.

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**He SAI**

A chemical equation is the symbol in Chemistry that represents chemical reaction with the help of chemical formulas. It contains the chemical substances that are involved in the reaction. It contains reactants and products. The reactants are the elements that react with one and another in a chemical reaction,

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while the products are  
the elements...

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+ OH-(aq) Equilibrium  
Shift none Same 10,6  
kcal Amount NaOH(s)  
Same Same 82

decreases same Same  
remains the increases  
same same same  
Same Same same  
same same

(Remember that pure  
solids and liquids do  
not affect equilibrium  
values.) [Na+] same Sa  
[OH-] same Same same  
Sa me same same  
sanqe



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